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Pearson Education Chapter 8 Covalent

Chapter 8 Covalent Bonding 183 Section Review Objectives • State a rule that usually tells how many electrons are shared to form a covalent bond • Describe how electron dot formulas are used • Predict when two atoms are likely to be joined by a double or a triple covalent bond • Distinguish between a single covalent bond and other covalent bonds

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Chapter 8

Michelle_McGee74 TEACHER. Pearson Chemistry Chapter 8: Covalent Bonding. Covalent Bond. Molecule. Diatomic molecule. Molecular compound. A bond formed by the sharing of electrons between atoms. A neutral group of atoms joined together by covalent bonds. A molecule consisting of two atoms.

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Pearson Chemistry Chapter 8 Vocabulary. Covalent bond. Molecule. Diatomic molecule. Molecular compound. bond formed by the sharing of electrons between atoms. a neutral group of atoms joined together by covalent bonds. a molecule that contains two atoms. a compound composed of molecules.

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Bookmark File PDF Chapter 8 Covalent Bonding Packet Answers Pearson Education configuration. In covalent bonding, atoms share electrons to achieve noble gas configuration. Most atoms share electrons until they have a total of 8 valence electrons (octet rule). Chapter 8 - Covalent Bonding covalent bonds. The majority

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8 Copyright © Pearson Education, Inc., or its affiliates. All Rights Reserved. In general, covalent bonding results from an imbalance between the attractions and repulsions of the nuclei and electrons involved. • The nuclei and electrons attract each other. • Nuclei repel other nuclei.

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