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Answers: #4 to end (quantitative  
problems, primarily) 4a) Balance would  
tip to the right:  $m_A > m_B$  for the same  
volume b)  $D_A = 60\text{g}/45\text{mL} = 1.3\text{g/mL}$ ;  $D_B = 32\text{g}/60\text{mL} = 0.53\text{g/mL}$  Each mL of  
substance has a mass of (1.3 or 0.53)  
grams. c)  $m_A = m_B$  when it is  
balanced.  $m_A = 1.3\text{ g/mL} * 10.0\text{ mL} =$

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13 g A => m = 13 g

## **Figure 1 B FIGURE 1 A B CP Chemistry Unit 1 Worksheet 3**

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Dimensional Analysis. Use the factor-label method to make the following conversions. Remember to use the appropriate number of sf's in your answer. Part 1 . 1. 74 cm x = meters. 2.

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8.32 x 10<sup>-2</sup> kg x = grams. 3. 55.5 mL x  
= cm<sup>3</sup>. 4. 0.00527 cal x = kilocalories

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